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### CHAPTER 2

## Atoms, Molecules, and Ions

### ■ SOLUTIONS TO EXERCISES

*Note on significant figures:* If the final answer to a solution needs to be rounded off, it is given first with one nonsignificant figure, and the last significant figure is underlined. The final answer is then rounded to the correct number of significant figures. In multistep problems, intermediate answers are given with at least one nonsignificant figure; however, only the final answer has been rounded off.

- 2.1. The element with atomic number 17 (the number of protons in the nucleus) is chlorine, symbol Cl. The mass number is  $17 + 18 = 35$ . The symbol is  $^{35}_{17}\text{Cl}$ .
- 2.2. Multiply each isotopic mass by its fractional abundance; then sum:
- $$\begin{array}{rcl} 34.96885 \text{ amu} \times 0.75771 & = & 26.496247 \\ 36.96590 \text{ amu} \times 0.24229 & = & \underline{8.956467} \\ 35.452714 & = & 35.453 \text{ amu} \end{array}$$
- The atomic mass of chlorine is 35.453 amu.
- 2.3. a. Se; Group VIA, Period 4; nonmetal  
b. Cs; Group IA, Period 6; metal  
c. Fe; Group VIIIB, Period 4; metal  
d. Cu; Group IB, Period 4; metal  
e. Br; Group VIIA, Period 4; nonmetal
- 2.4. Take as many cations as there are units of charge on the anion and as many anions as there are units of charge on the cation. Two  $\text{K}^+$  ions have a total charge of  $2+$ , and one  $\text{CrO}_4^{2-}$  ion has a charge of  $2-$ , giving a net charge of zero. The simplest ratio of  $\text{K}^+$  to  $\text{CrO}_4^{2-}$  is 2:1, and the formula is  $\text{K}_2\text{CrO}_4$ .

# General Chemistry Ebbing Gammon Instructor Solution Manual

**David Bookin, Darrell D. Ebbing, Steven  
D. Gammon**



## **General Chemistry Ebbing Gammon Instructor Solution Manual:**

*Student Solutions Manual for Ebbing/Gammon's General Chemistry, 11th* Darrell Ebbing, Steven D. Gammon, 2016-02-09  
This manual contains answers and detailed solutions to all the in chapter Exercises Concept Checks and Self Assessment and Review Questions plus step by step solutions to selected odd numbered end of chapter problems **General Chemistry**  
Darrell Ebbing, Steven D. Gammon, 2005 Known for its carefully developed thoroughly integrated approach to problem solving this market leading text emphasizes the conceptual understanding and visualization skills essential for first year chemistry and science majors The new technology program reinforces the approach of the text and provides a complete solution for teaching and learning The Eighth Edition retains the hallmark pedagogical features of the text and builds upon the conceptual focus and art program Students also benefit from online homework in the technology program which features an extensive database of questions drawn from the text In order to reinforce major chemical concepts the authors present a proven six part approach to problem solving that includes Example Problem Strategy Solution Answer Check Exercise and corresponding End of Chapter Problems many of which are presented in matched pairs The Media Integration Guide for Instructors includes several user friendly supplements designed to make class preparation presentation and course management more efficient and effective HM ClassPrep HM Testing CD ROM with images customizable test bank instructor resource manual and solutions manual HM ClassPresent CD ROM with animations organized by topic instructor web site access and information about Eduspace powered by Blackboard Eduspace is Houghton Mifflin s online learning tool Powered by Blackboard Eduspace is a customizable powerful and interactive platform that provides instructors with text specific online courses and content This Ebbing et al General Chemistry course features test bank material for exams algorithmic in chapter and dynamically generated end of chapter homework problems from question pools and access to ACE quiz content for independent practice For the instructor we also provide presentation slides photos illustrations interactive tables and video clips The Media Guide for Students provides information on and access to multimedia tools that help students visualize chemical concepts and practice problem solving strategies SMARTHINKING live online tutoring student web site with animations and Student CD ROM The guide also includes information about Eduspace powered by Blackboard

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